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Determine the molecular weight and stoichiometric mole and mass air-fuel ratios for the Oklahoma gas mole composition given in Table 3.4. Solution. Equation (3.2), $CH_4 + 2O_2 +$

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$2(3.76)N_2CO_2 + 2H_2O + 2(3.76)N_2(3.2)$ shows that there are $2 + 2(3.76) = 9.52$ moles of air required for complete combustion of each mole of methane.

FUELS AND COMBUSTION 3.1 Introduction to Combustion

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This solutions manual has been prepared to accompany the 3rd edition of the author's Introduction to Internal Combustion Engines. At the end of many of the questions is a discussion, which is intended to provide useful supplementary information.

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