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296 CHAPTER 12 CHEMICAL KINETICS

$2.30 \times 10^{-1} = k(0.100)(0.100)$ y and $1.15 \times 10 = k(0.100)(0.0500)$ Dividing: $2.00 = 2.00y$, $y = 1$ The rate law is: $\text{Rate} = k[\text{ClO}_2]^2[\text{OH}^-]$
 $2.30 \times 10^{-1} \text{ mol/LCs} = k(0.100 \text{ mol/L})^2(0.100 \text{ mol/L})$, $k = 2.30 \times 10 \text{ L/mol Cs} = k \text{ mean b. Rate} = k[\text{ClO}_2]^2[\text{OH}^-] = 0.594 \text{ mol/LCs}$ Integrated Rate Laws 27.

~~CHAPTER TWELVE CHEMICAL KINETICS~~

Chapter 12: Chemical Kinetics. chemical kinetics. thermodynamic favorability. Factors that affect reaction rates. nature of the reactants. the study of the speed or rate of a reaction under various con.... the energy state of reactants is higher than that of the produ.... 1. nature of the reactants... 2.

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Chapter 12 - Chemical Kinetics - Review
Questions - Page 591: 1. Answer. Reaction
rate: rate at which the concentration of a
reactant or product changes over time
Initial Rate: reaction rate at the instant the reaction
begins
Average Rate: reaction rate over an
interval of time
Instantaneous rate: reaction
rate at an instant in time
The initial rate is
usually the fastest.

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Chemical Kinetics ...~~

Chapter 12 - Chemical Kinetics . 12.1
Reaction Rates . A. Chemical kinetics 1.
Study of the speed with which reactants are
converted to products
B. Reaction Rate 1.
The change in concentration of a reactant or
product per unit of time $[] t A t t$
concentration of A at time t concentration

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of A at time t Rate = $-\frac{1}{2} \frac{d[A]}{dt} = 2 \frac{1}{2} \frac{1}{t}$. a.
Rates decrease with time b.

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Questions Chemical Kinetics Class 12

Important Questions Very Short Answer

Type Question 1 Define ‘ rate of a

reaction ’ (Delhi 2010) Answer: Rate of a

reaction: Answers Chapter 4 Chemical

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Chemistry Chapter 4 Chemical Kinetics:

4.1. For the reaction $R \rightarrow P$, the

concentration of reactant changes from 0.03

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M to 0.02 M in 25 minutes. Calculate the average rate of reaction using units of time both in minutes and seconds.

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the coefficients in the balanced equation: t

$[H_2O]_{21} - t [O_2] = - 2 2 = 1.16 \times$
 $10^{-5} \text{ mol/LCs b. } (4.32 \ 10 \ 2.16 \ 10) \text{ s } (0.250$

$0.500) t [H_2O]_{44} 2 2 \times - \times - -$

$- = M = 1.16 \times 10^{-5} \text{ mol/LCs} - t$

$[O_2]$ CHAPTER 12 CHEMICAL
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Chemical Kinetics Class 12 Chemistry MCQs Pdf. 1. The half life period of first order reaction is 1386 seconds. The specific

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rate constant of the reaction is (a) $0.5 \times 10^{-2} \text{ s}^{-1}$ (b) $0.5 \times 10^{-3} \text{ s}^{-1}$ (c) $5.0 \times 10^{-2} \text{ s}^{-1}$ (d) $5.0 \times 10^{-3} \text{ s}^{-1}$.

Answer/Explanation. Answer: b

Explanation:

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Chemical Kinetics Class 12 MCQs

Questions with Answers. Question 1. In

chemical equation $\text{H}_2(\text{g}) + \text{I}_2(\text{g})$

\rightleftharpoons $2\text{HI}(\text{g})$ the

equilibrium constant K_p depends on (a)

total pressure (b) catalyst used (c) amount of H_2 and I_2 (d) temperature. Answer.

Answer: (b) catalyst used

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For students of class 12, it is important that they are clear on every topic of chemistry.

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Answer Type Question 1. Define ' rate of a
reaction ' . (Delhi 2010) Answer: Rate of a
reaction: Either, The change in the
concentration of any one of the reactants or
products per unit time [...]

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Mark Questions and Answers. Question 1.

Explain a graphical method for determination of activation energy. Answer: Activation energy can be determined graphically from the $\ln k$ vs $1/T$ graph. From the graph, $\ln k = \ln(Ae^{-E_a/RT})$ $\ln k = \ln A + \ln e^{-E_a/RT}$ $\ln k = \ln A + -E_a/RT$ This is in the form of $y = mx + c$

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1. The rate of a chemical reaction tells us about. the reactants taking part in the reaction; the products formed in the reaction; how slow or fast the reaction is taking place; none of the above; Answer: (c)
2. In the rate equation, when the concentration of reactants is unity then the rate is equal to . specific rate constant; average rate constant

~~MCQ on Chemical Kinetics for NEET 2020~~

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Chemical Kinetics Answers: (a) 8.4×10^{-7} M/s, (b) 2.1×10^{-7} M/s
SAMPLE EXERCISE 14.3 continued The decomposition of N_2O_5 proceeds according to the following equation: If the rate of decomposition of N_2O_5 at a particular instant in a reaction vessel is 4.2×10^{-7} M/s, what is the rate of appearance of (a) NO_2 , (b) O_2 ?

~~Chapter 14 Chemical Kinetics—University of Massachusetts ...~~

A1: The various concepts, topics, and subtopics that students can revise from the class 12 chemistry notes chapter 4 chemical kinetics are as mentioned below: 4.1 The rate of a Chemical Reaction. 4.2 Factors Influencing the Rate of a Reaction. Dependence of Rate on Concentration. Rate Expression and Rate Constant. Order of a Reaction

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