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Solution This problem gives us, for a hypothetical ion pair, values for  $r_0$  (0.35 nm),  $E_0$  (6.13 eV), and  $n$  (10), and asks that we determine explicit expressions for attractive and repulsive energies of Equations 2.8 and 2.9. In essence, it is necessary to compute the values of A and B in these equations.

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Solution This problem is solved using two steps: (1) calculate the total number of lattice sites in silver,  $N_{Ag}$ , using Equation 4.2; and (2) multiply this number by fraction of lattice that are vacant,  $2 \cdot 10^{-6}$ . The parameter  $N_{Ag}$  is related to the density,  $(\rho)$ , Avogadro's number ( $N_A$ ), and the atomic weight ( $A_{Ag} = 107.87$  g/mol, from inside the

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Solution The average atomic weight of silicon ( $A_{Si}$ ) is computed by adding fraction-of-occurrence/atomic weight Si products for the three isotopes—i.e., using Equation 2.2. (Remember: fraction of occurrence is equal to the percent of occurrence divided by 100.) Thus.  $A_{Si} = f_{28} A_{28} + f_{29} A_{29} + f_{30} A_{30}$

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Adapted from Fig. 9.3(a), Callister 7e. (Fig. 9.3(a) is adapted from Phase Diagrams of Binary Nickel Alloys , P. Nash (Ed.), ASM International, Materials Park, OH (1991). • 2 phases: L (liquid)  $\alpha$  (FCC solid solution) • 3 phase fields: L L +  $\alpha$   $\alpha$  0 20 40 60 80 100 wt% Ni 1000 1100 1200 1300 1400 1500 1600 T(°C) L (liquid)  $\alpha$  (FCC solid ...

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Chapter 9: Phase Diagrams

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Tools to reframe the solution -> problem statement. This kind of question may be much easier for researchers to think about since we are used to doing this daily. For our colleagues, it might not come as naturally. Luckily, there are a few different methods we can employ to help us reframe the solution.

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