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solids are dense and not easy to compress. Because the particles in solids tend to vibrate about fixed points, solids do not flow. When you heat a solid, its particles vibrate more rapidly as their kinetic energy increases. The organization of particles within the solid breaks down, and eventually the solid melts.

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We can make all solids liquids, though, by increasing the internal kinetic energy to where it becomes stronger than the intermolecular forces keeping it as a solid.

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The general properties of solids reflect the orderly arrangement of their particles and the fixed locations of their particles.

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Some solids lack an ordered internal structure and are called amorphous solids. In particular, when the atoms in a solid have a random arrangement, the solid is a glass.

13.3 section summary.doc - SECTION SUMMARY 13.3 The Nature ...

Home □ Characteristics and Nature of Solids In solid state, the constituent particles are closely packed and the voids between them are very small.

Characteristics and Nature of Solids | Chemistry Assignment

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PowerPoint Presentation

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Constituent particles in ionic solids of the Crystalline Solids are anions (negatively charged) and cations (positively charged). An ion is surrounded by a typical number of opposite charges. For

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example, in NaCl, the Na⁺ ion is surrounded by 6 Cl⁻ ions. Ions in these solids are held together by strong electrostatic forces.

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